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12.1 The Arithmetic of Equations • A balanced chemical equation provides the same kind of quantitative information that a ... Stoichiometry 379 CHAPTER 12 Assessment 36. a. Two formula units KClO_3 decom-pose to form two formula units KCl and three molecules O_2 . b.

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stoichiometry. 2. What is always conserved in every chemical reaction? 12.2 Chemical Calculations. 1. Define mole ratio. 2.

In a balanced chemical equation, where can one find the numbers to make up the mole ratio? 3.

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Pharisees followed Him making blasphemous accusations

against the Spirit and demanded a sign (22-45). Matthew 12

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1. Define the following: a. Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.

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luis_dewendt. Terms in this set (14) Avogadro's Number. A constant that represents the number of constituent particles (usually atoms or molecules) per mole of a given substance (6.02×10^{23})

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... In a typical Stoichiometry problem, the given quantity is first converted to _____. Then, the mole ratio from the balanced

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equation is used to calculate the moles of unwanted substance. ... Chem Quiz 12.1-12.3. 20 terms ...

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, 154.21 g/mol ($6 \times 12 + 5 \times 1 = 77$: $154/77 = 2$) C₁₂ H₁₀ C₃ H₆ O₂, 222.24 g/mol ($12 \times 3 + 6 \times 1 + 2 \times 16 = 77$: $222/74 = 3$) C₉ H₁₈ O₆ 3. Empirical Formula from Percentage Composition (don't round until the very end!!!!) A compound is composed of 12.590% hydrogen, 37.404% aluminum, and 50.006% carbon. Please find the empirical formula:

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