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Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

1.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$  Sr 0 to  $\text{Sr}^{2+}$ ;  
oxidized/reducing agent  $\text{O}_0$  to  $\text{O}^{2-}$ ;  
reduced/ox. ag.
2.  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$  Li 0 to  $\text{Li}^+$ ;  
oxidized/red. ag. S 0 to  $\text{S}^{2-}$ ; reduced/ox. ag.
- 3.

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Chapter 20 Worksheet: Redox I. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also. 1.

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Chapter 20 Worksheet: Redox ANSWERS I. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

1. 2. 3. 4. 5. 6.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$  Sr 0 to  $\text{Sr}^{2+}$  ;  
oxidized/reducing agent  $\text{O}_0$  to  $\text{O}^{2-}$ ;  
reduced/ox. ag. Li to  $\text{Li}^+$  oxidized/red, ag. so to reduced/ox. ag. N 0 to  $\text{N}^{3-}$ ; reduced/ox. ag,  $(.)_0$  to  $(l)_1$  I w; reduced/ox. ag. Br to  $\text{Br}^-$   
oxidized/red. ag.

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1. The oxidation number of any uncombined element is 0.
2. The oxidation number of a monatomic ion equals the charge on the ion.
3. The more-electronegative

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1  $2\text{Sr} + \text{O}_2$  2  $2\text{SrO}$  2  $2\text{Li} + \text{S}$   $\text{Li}_2\text{S}$  3  $2\text{Cs} + \text{Br}_2$   $2\text{CsBr}$  4  $3\text{Mg} + \text{N}_2$   $\text{Mg}_3\text{N}_2$  5  $4\text{Fe} + 3\text{O}_2$   $2\text{Fe}_2\text{O}_3$  6  $\text{Cl}_2 +$

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Determine what is oxidized and what is  
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1.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$  Sr 0 to  $\text{Sr}^{2+}$ ;  
oxidized/reducing agent  $\text{O}_2$  to  $\text{O}^{2-}$ ;  
reduced/ox. ag. 2.  $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$  Li 0 to  
 $\text{Li}^{+}$ ; oxidized/red. ag.  $\text{S}^0$  to  $\text{S}^{2-}$ ;  
reduced/ox. ag. 3.  $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$

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